

ACIDS and BASES

Different Types of Acids + Bases :

- Arrhenius Acid: creates H^+ in H_2O
- Arrhenius Base: creates OH^- in H_2O
- Bronsted-Lowry Acid: H^+ donor
- Bronsted-Lowry Base: H^+ receiver
- Lewis Acid: accepts a pair of electrons
- Lewis Base: donates a pair of electrons

